

Binary Covalent Compounds

<p>Step 1: The first element in the name and formula of a binary covalent compound is usually the one that is farthest to the left on the periodic table</p>	<p>CO <i>carbon monoxide</i></p> <p>Carbon comes first because carbon is to the left of oxygen on the periodic table</p>
<p>Step 2: When naming, the suffix -ide is attached to the name of the second element</p>	<p>CO <i>carbon monoxide</i></p> <p>Oxygen is changed to oxide in the name carbon monoxide</p>
<p>Step 3: When naming, prefixes are used to indicate how many atoms of each type are present in one molecule of the compound</p>	<p>CO <i>carbon monoxide</i></p> <p style="text-align: right;">mono means one</p> <p>When there is no prefix, mono is implied</p>
	<p>When mono comes before oxide, an "o" is dropped</p>
	<p>When the addition of a prefix results in two vowels together, the vowel at the end of the prefix is usually dropped....</p> <p>Except for prefixes di- and tri- The "i" at the end are never dropped</p> <p>PI₃ phosphorus triiodide</p>

Prefix	Number	Prefix	Number
mono-	1	hexa-	6
di-	2	hepta-	7
tri-	3	octa-	8
tetra	4	nona-	9
penta-	5	deca-	10

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Write the formulas for each of the following covalent compounds

sulfur tetrafluoride _____

disulfur difluoride _____

dinitrogen trioxide _____

oxygen difluoride _____

nitrogen tribromide _____

diiodine hexachloride _____

Write the names of the following covalent compounds

PI_3 _____

SO_2 _____

SO_3 _____

S_2F_{10} _____

CCl_4 _____

N_2O_5 _____

N_2O _____

NI_3 _____

P_2O_5 _____

PBr_5 _____

As_2S_3 _____

ICl_3 _____